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WAEC Mathematics Questions 2020/2021 Objective and Theory ...

We can use the Arrhenius equation to relate the activation energy and the rate constant, k , of a given reaction:

$(k=A\{e\}^{\{-\}E_{\{a\}}\}/RT})$ In this equation, R is the ideal gas constant, which has a value 8.314 J/mol/K , T is temperature on the Kelvin scale, E_a is the activation energy in joules per mole, e is the constant 2.7183 , and A is a constant called the frequency ...

Activation Energy and the Arrhenius Equation | Chemical

...

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Elective Mathematics, Elective chemistry), Technical (Elective Physics, Elect ive Chemistry, Elective Mathematics, Technical Drawing, Wood Carving, Metal Works, Building Techno logy) [22].

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